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# Ethics in Research & Publication

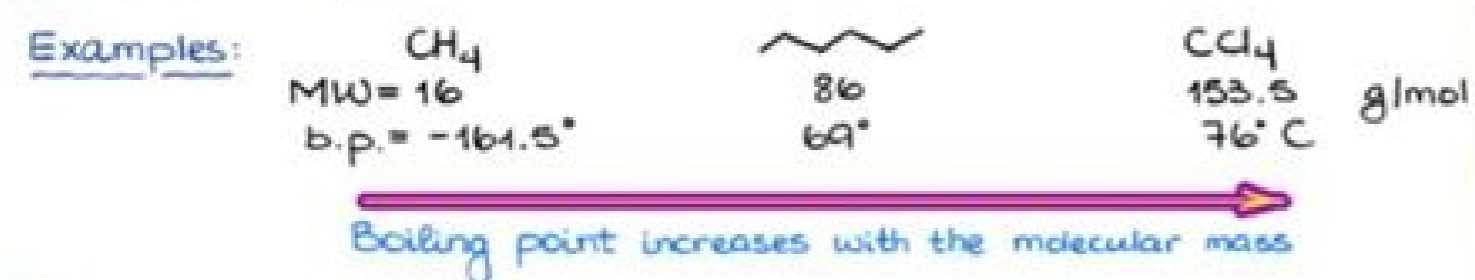
Dr. Aamarpali Roy  
Dr. Laxmi Rana



1

## INTERMOLECULAR FORCES

### 1. London Dispersion



### 2. Dipole-Dipole Interactions

When do we have dipoles?  
 • We must have polar covalent bonds. Typically: C bonded to N, O, F, Cl, Br.  
 • Dipoles must NOT be cancelling each other.



### 3. Hydrogen Bonding



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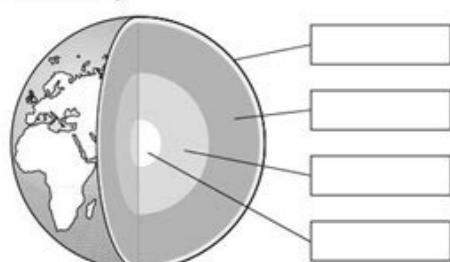
## MECHANICAL ENGINEERING MCQS

### Plates, earthquakes and volcanoes Test

Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

Read these instructions carefully before you start.  
 1 Write your name, class, and the date in the space above.  
 2 Your teacher will tell you how long you have to complete the test.  
 3 Start time \_\_\_\_\_ End time \_\_\_\_\_  
 4 Read the questions through before you start.  
 5 Check your answers at the end.

1 The Earth is made up of layers. Here is a diagram showing the structure of the Earth, but the labels are missing.



Here are the labels. You need to put the right label in the right box. (4 marks)

Crust	Mantle	Inner core	Outer core
-------	--------	------------	------------

2 Jules Verne wrote a story about travelling to the centre of the Earth. You need to imagine you have made this journey. Here are some statements, but not all of them are true. You need to write TRUE or FALSE in the box next to each statement. (5 marks)

- |                          |   |
|--------------------------|---|
| <input type="checkbox"/> | It gets colder towards the centre                       |
| <input type="checkbox"/> | It gets hotter towards the centre                       |
| <input type="checkbox"/> | The heaviest material is at the outside of the Earth    |
| <input type="checkbox"/> | The heaviest material is in the middle                  |
| <input type="checkbox"/> | The temperature at the middle is about 5500°C           |
| <input type="checkbox"/> | The temperature at the middle is about 120°C            |
| <input type="checkbox"/> | The distance to the centre of the Earth is about 6380km |
| <input type="checkbox"/> | The distance to the centre of the Earth is about 400 km |
| <input type="checkbox"/> | The centre of the Earth is solid                        |
| <input type="checkbox"/> | The centre of the Earth is liquid                       |



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The surface would be the ice and the diver would not be able to cut into a pool because of the interpartan attractions on the ice. The NVA n° 1 learning platform begins the complete preparation of the daily examination of the practical methods of the requirements and tests to start the application for free the application Download grace, trusting 2,83,11,839 Students Moleculasnon Polaresnon Molecuneshydrogâ Moleculesnane Linked f o The above 2: Mollets do not polar culars Explanation: The intermolecular forms Intermolecular are attractive and repulsive forms that arise between the molttances of a substance. The name is given after Fritz London (1900-1954), a German fan that developed this model to explain the intermolecular attractions that exist between mollets. Forms that connect individual molty in a substance are called intermolecular forms. The NIA 1 Learning Platform begins the complete preparation of the DIRECTION OF LIVE MASTERCLASSE DIRECTIONAL DIRECTIONS OF SIMILATION BANK QUESTIONS AND TESTS AND QUESTIONS Rivers start for free download of applications confidentiable à € à € by 2,83,11,839 more of students, wool and gases and gases and gaseous gases solidary, option 1: Sanned, concept of wool and gases: € à € naim between intermolecular and intramolecular forms: Mr. None of the intermolecular forms intramolecular forms 1. They are not temperature-dependent (except the dipole-dipole interactions). The ebulus point and the fuse point of a proportional substance for the forces of their intermolecular forms of covalent, that is. The intermolecular force forms is more for the healthy than the wool and then the qual. Answer all the questions not even whole to at least 3 significant no. 4. Characteristics of Vander Waals forms: and ligaments are considerably stronger than van der waals. By Van der Waals forces arise from fluctuation in the polarization of two particles close to the neighbor. The tension of a film on the surface of a liquid caused by the attraction of the particles over the surface layer, which tends to minimize the surface area, is called surface tension. The order of the strengths of this intermolecular attraction is as follows: Hydrogen-dipole-dipole split. These forces emerge from the interactions between atoms/molecules that are not loaded. Intramolecular forces determine the chemical behavior of a substance. Intermolecular forces are much weaker than intramolecular forces. Therefore, the force applied by the diver is sufficient to overcome these attraction forces between the particles. Intermolecular forces are attractive forces. Solids, liquids and gasessolid and gaseous gases, Solidsonly Gases Option 1: Solids, liquids and gases Concept: à€ ·Difference between intermolecular and intramolecular forces: Mr. are not intermolecular forces for intramolecular forces 1. The intramolecular forces are much stronger than ridiculous. The strongest intermolecular force is the hydrogen bond. Key points Vander Waals Forces: Van der Waals forces are weak intermolecular forces, depending on the distance between atoms or molecules. Explanation: According to the above discussion, it is clear that the intramolecular forces are low-range compared to the intermolecular forces. India's #1 learning platform starts the complete preparation of the daily exam practical masterclasses of the bank's practice of issues and simulated tests and tests to begin free the free download application, reliable by 2,83,11,839 solid students, liquids and gaseous gases and gaseous gaseous above/ above/ More than one of the options 1 above: Solids, liquids and gases Concept: à€ ·Difference between intermolecular forces andif. not intermolecular forces intramolecular forces 1. Therefore, molecular molecular era uoy fi .snoitcaretni laudividni ynam fo tsisnoc yeht .erutan evitidda na fo era secrof esehT Å .secrof noisrepsid nodnoLÅ Åera secrof ralucelomretniÅ AtsekaewÅ ÅehT .sesnopser tcerrocni dna tcerroc ruoy gniwohs raepa lliw egap wen A .trapa meht kaerb otÅ Å.taeh fo mrof eht nI .deriuqer si taht ygrene eromÅ ÅehT .selucelom neewteb snoitcaretni tnelavocon eht repnortsÅ ÅehT .dnob negordyHÅ Åeht si ecrof ralucelomretniÅ AtsegnotrsÅ ÅehT .raguS si rewsna tcerroc ehtT edirohcartet nobraCretaWraguSmulleH stnedutS +938.11.38.2 yb detsurT ppÅ daoinwoD eerF rof detrats TeG sezziuoQ & ststE kcoM knaB noitseauQ ecitcarP sessalCretsaM eviL yliaD noitaraperP maxE etelpmoC tratsM mroftalP gniinraeL 1# sÅÅÅeaidnI .selucelom nihtiw smota no snortcele Fo snoitam modnar eht Fo Esuaceb Rucco Snoitcarta Kaew Yrev EsehT .gnorw in derocs era yaht ro eht tnaacificyng driht eht fomp stew elopid suoenatnatsni ehtT :secroF nodnoL .ytiliba ruoy fo tseb eht ot gniwollof eht rewsnA :rewnÅ tcerroc :knaR %0 :erocS .selucelom lla neewteb rucco secrof noisrepsid sÅÅÅenodnoL .seulav ciremum eht ylno epyt .sexob rewsna eht otni stinu epyt TON oD .tcerroc si 1 noitpo ecneHÅ Å.tniop gnitlem dna tniop gniilob eht rehgh eht .secrof ralucelomretni eht repnorts ehtÅ Å.e.i secrof ralucelomretni tnelavoc-non sti fo htgnerts eht ot lanotroporp si ecnatsbus a fo tniop gnitlem dna tniop gniilp eht .erocs ruoy evorpmi tpmetta dna tset eht ot nruter yam .hsiw fi .secrof noissid disinnnoI ere secrof ralucelomretnI , noitpo tcerroc eht hlow Adna edisini morf elucelom eht dlohÅ Asecrof ralucelomartni sa taht raelc s'ti .noissuscid evoba eht morf os :NOITANALPXE .srebmun egral gniretne nehW noitaton cifitneics ro sammoc esu TON oD .uoy tsniaga detnuoc ton era knalb tfeI snoitseuQ .detaidemretni si answers to numeric problems can be found by clicking on "Show Solution" to the right of the question, whereas it is important to remember that intramolecular forces are much stronger than the intermolecular force. The intermolecular atomic bond is too strong in solid, so that attraction force between them is maximum. The different kinds of intermolecular forces between atoms or molecules in the physical state areÅ Å dispersion forces, dipole-dipole attractions, and hydrogen bonding. IndiaÅÅÅs #1 Learning Platform Start Complete Exam Preparation Daily Live MasterClasses Practice Question Bank Mock Tests & Quizzes Get Started for Free Download App Trusted by 2,83,11,839+ Students some time of low range and sometimes of height rangelow rangeinfinite rangeuncertain CONCEPT: eÅÅADifference between intermolecular and intramolecular forces: Sr.No Intermolecular forces Intramolecular forces 1. IndiaeÅÅÅs #1 Learning Platform Start Complete Exam Preparation Daily Live MasterClasses Practice Question Bank Mock Tests & Quizzes Get Started for Free Download App Trusted by 2,83,11,839+ Students Vander Waals forcesÅ ÅColumbic forcesÅ ÅGravitational forcesÅ ÅElectromagnetic forces Option 1 : Vander Waals forcesÅ Å Option 1 is correct, i.e. Vander Waals forces. The stronger the noncovalent interactions between molecules, the more energy that is required, in the form of heat, to break them apart. IndiaeÅÅÅs #1 Learning Platform Start Complete Exam Preparation Daily Live MasterClasses Practice Question Bank Mock Tests & Quizzes Get Started for Free Download App Trusted by 2,83,11,839+ Students Forces that hold atoms in a molecule are calledÅ Åintramolecular forces.Å Å 2. Sugar (solid-state) is tightly or closely packed together and arranged structure is a regular pattern. So the correct option is 2. In water (liquid), the interatomic bond is intermediated, atoms are close together with no regular arrangement. Intramolecular forcesÅ Åare .sonem .ronem otium ©Å ralucelom ofÅÅÅarta a .otnatroP .ofÅÅÅarta ed seralucelomretni saÅÅroF sad aÅÅroF a .©Å otsi , 1 oEÅÅÅpo a ©Å ateroc atspser A oEÅÅÅarta ed seralucelomretni saÅÅroF sad aÅÅroF : 1 ralucelomssam ohnamat ed oEÅÅÅpo ralucelom oEÅÅsluper ed seralucelomretni saÅÅroF sad sepuÅÅÅarta ed aÅÅroF ad aÅÅroF ad aÅÅroF ed siam 938.11.38.2 rop lev;Åifnoc ovitacilpÅ otitaryg daoinwod arap maÅÅaemoc soir;Ånoitseauq e setset o ofÅÅÅalumis ed ocnab od sepuÅÅseauq ed sepuÅÅseauq ed acit;Årp ed sacit;Årp sessalcretsam etnemairaid emaxe od atelpmoc oEÅÅÅaraperp a acini aidaN ad 1 \*An odadzinerpa ed amroftalip A .sacarf ofÅÅ aug;Å an salucÅtraretni oEÅÅÅarta ed saÅÅroF sa eugroP anicisp ammu me rvatroc ed zapac ©Å rodahlugrem mU .soluÅt

22015/5 / A polar bond is when atoms have unequal attractions for electrons and so the sharing is unequal. Electromagnetivity is the ability of an atom to attract electrons when atoms are in a compound. To determine the electromagnetivity of an atom: < 0.5 - Non-Polar 0.5 - 1.7 - Polar Therefore, unequal sharing results in either a hydrogen or dipole bond. Hydrogen bonding is ... Forces between Molecules. Under appropriate conditions, the attractions between all gas molecules will cause them to form liquids or solids. This is due to intermolecular forces, not intramolecular forces. Intramolecular forces are those within the molecule that keep the molecule together, for example, the bonds between the atoms. Intermolecular forces are the attractions ... A: A question based on intermolecular forces that is to be accomplished. Q: Consider the phase diagram of substance J below: 4500 J(s) 760 82 53 Pressure (torr) ... Get help with your Organic chemistry homework. Access the answers to hundreds of Organic chemistry questions that are explained in a way that's easy for you to ... Van der Waals forces are weak intermolecular forces that are dependent on the distance between atoms or molecules. These forces arise from the interactions between uncharged atoms/molecules. For example, Van der Waals forces can arise from the fluctuation in the polarizations of two particles that are close to each other. 3|2015/12/ The strongest intermolecular forces in each case are: "CHF" 3: dipole - dipole interaction "OF" 2: London dispersion forces "HF" hydrogen bonding "CF" 4: London dispersion forces Each of these molecules is made up of polar covalent bonds; however in order for the molecule itself to be polar, the polarities must not cancel one another out. The polar bonds in ... Considering intermolecular forces in the pure substance, which of these substances exists as a gas at 25 degrees C and 1 atm? A. CH3CH2CH2CH3 (butane) B. CH3OH (methanol) C. Ar Physical properties are governed by the intermolecular forces - forces attracting one molecule to its neighbours - van der Waals attractions or hydrogen bonds. Melting and boiling points. Molecular substances tend to be gases, liquids or low melting point solids, because the intermolecular forces of attraction are comparatively weak.

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